Reaction Rates (MCQ)

1. What is the **main** reason for the increase in reaction rate with increasing temperature?

A The activation energy decreases.
B The activation energy increases.
C More molecules have an energy greater than the activation energy.
D The molecules collide more frequently.

Your answer ☐ **[1]**

2. Which statement explains why the rate of a reaction increases when the temperature is increased?

A. The activation energy for the reaction decreases.
B. The activation energy for the reaction increases.
C. The proportion of molecules exceeding the activation energy decreases.
D. The proportion of molecules exceeding the activation energy increases.

Your answer ☐

[1]

END OF QUESTION PAPER

Mark scheme – Reaction Rates (MCQ)

Question			Answer/Indicative content	Marks	Guidance
1			**C**	1 (AO1.1)	**Examiner's Comments** The role of activation energy in the rate of a reaction with increasing temperature was well-known and most candidates chose the correct option C. From the annotations on candidate scripts, many had ruled out options A and B entirely. D was anticipated as being the main distractor and this proved to be the case. Activation energy has a much greater effect than increasing collision frequency.
			Total	**1**	
2			D	1	
			Total	**1**	